8.7 Review of Acid-Base Titration Lab Review

|  |  |
| --- | --- |
| Concept | Definition |
| Titration (p. 540) |  |
| Sample (p. 540) |  |
| Titrant (p. 540) |  |
| Burette (p. 540) |  |
| Standard Solution (p. 540) |  |
| Primary standard (p. 540) |  |
| Equivalence point (p. 540) |  |
| Endpoint (p. 540) |  |

8.8 Buffer Systems

1. Define buffer (for exam and assessment):

Review and Acid and Base pH buffer system:

2. Create a buffer system for an ethanoic acid.

3. Create a buffer system with a base buffer system:

4. Calculate an ICE table for the following acidic calculations; Ka for ethanoic acid is 1.8x 10-5 and O.50 mol/L for ethanoic acid and the ion 0.50 mol/L. Determine the pH of the system.

5. Calculate the pOH of a basic buffer system given the following:

Kb of ammonia is 1.8x10-5 and initial 0.25 mol/L; the initial concentration of the ion of ammonia is 0.40 mol/L. Determine the pOH of the sytem.

Review Questions p. 1-5.