Exam Revision

Balance the following equations

A)
$$\underline{\qquad}$$
 Li₃PO₄ + $\underline{\qquad}$ NaBr \rightarrow $\underline{\qquad}$ Na₃PO₄ + $\underline{\qquad}$ LiBr

B)
$$\subseteq$$
 CaF₂ + \subseteq Li₂SO₄ \rightarrow \subseteq CaSO₄ + \subseteq LiF

C)
$$\longrightarrow$$
 HBr + \longrightarrow Mg(OH)₂ \rightarrow \longrightarrow MgBr₂ + \longrightarrow H₂O

D) ____ LiNO₃ + ____ CaBr₂
$$\rightarrow$$
 ____ Ca(NO₃)₂ + ____ LiBr

Classify the following equations:

6) **2** CoBr₃ + **3** CaSO₄
$$\rightarrow$$
 3 CaBr₂ + **1** Co₂(SO₄)₃_____

Calculation questions:

7) How much 0.89 M of Na_3P is reguried in (mL) to react with 0.89 g of CaF_2 . Show all of your working out.

2 Na₃P + 3 CaF₂
$$\rightarrow$$
 6 NaF + 1 Ca₃P₂

8) If 0.89 g of LiO_3 and 2.89 g of $CaBr_2$ are used, determine which is the limiting reagent to produce LiBr. Show all working out.

1 **2** LiNO₃ + **1** CaBr₂ \rightarrow **1** Ca(NO₃)₂ + **2** LiBr

- 9. Determine what the definition of a titration is:

 10. Define equivalence point
- 11. Reflection on some of the issues that were discussed in class:
- 1. Is science always beneficial or always a hindrance
- 2. Has society always benefited from science invention?
- 3. How has technology changed or improved science discoveries?