Calculations Involving Basic Solutions

Chapter 8.5

Strong Bases

- A solution of calcium hydroxide has a concentration of 0.05 mol/L. Calculate the $[H^+]$, $[OH^-]$, and the pH of the solution.

Weak Bases

- A solution of aniline, $C_6H_5NH_2$, has a concentration of 5.0g/L and the pH of the solution is 8.68. Calculate the K_b for analine.

Weak Bases

- Quinine, $C_{20}H_{24}N_2O_2$, has a K_b of 3.3×10^{-6}. What are the hydroxide ion concentration and pH of a 3.6×10^{-3} mol/L solution of quinine?

HOMEWORK

Required Reading:

p. 526-530

(remember to supplement your notes!)

Questions:

p. 527 #1,2

p. 529 #1-3

p. 530 #1-10b